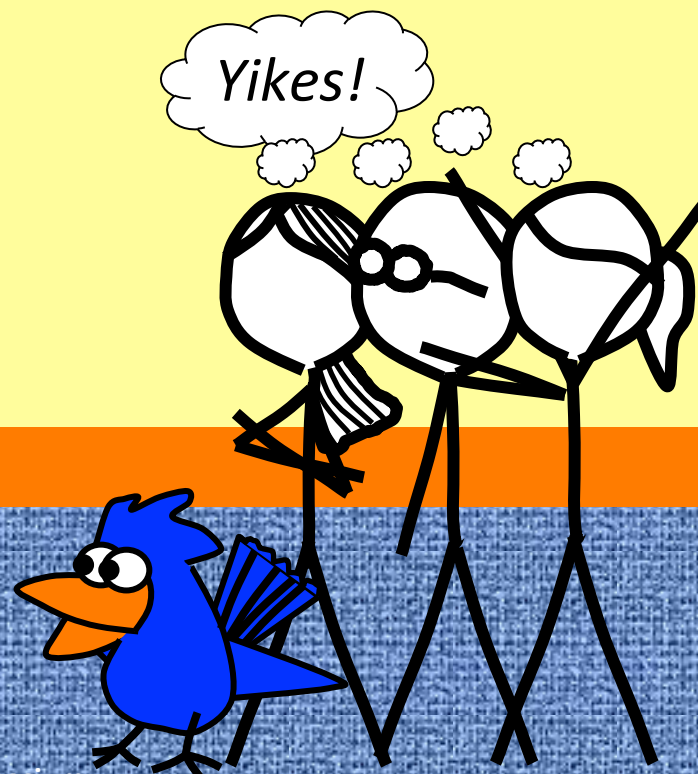


Experiment 5

28 September 2017

Classes of Chemical Reactions



<http://www.gettyimages.com/detail/photo/burning-magnesium-in-oxygen-high-res-stock-photography/128545618>

Objectives: To perform the types of reactions we have studied in Chapter 4.

Today we explore the sort of reactions we study in Chapter 4 in class.

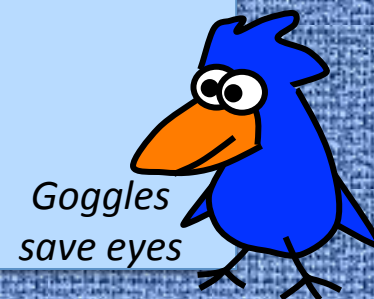
*All set here!
Let's go!*

Remember what I said last week about writing good introductions based on stuff from the presentation?

Overview:

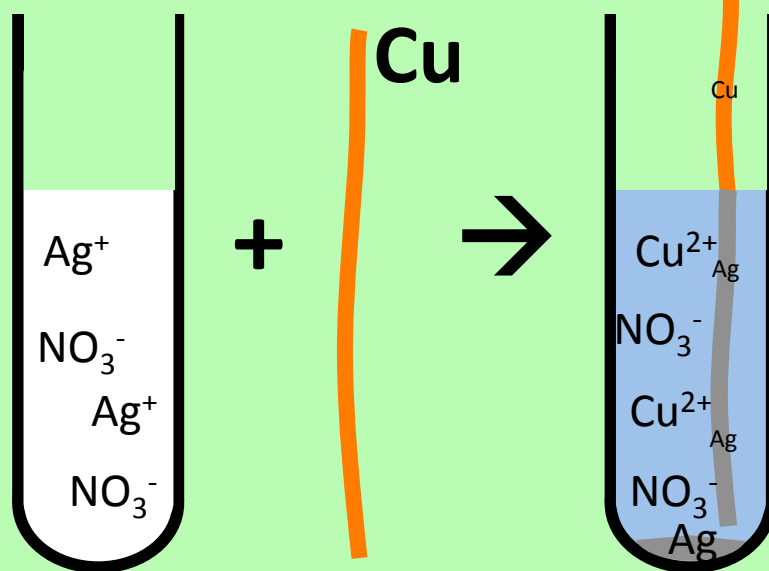
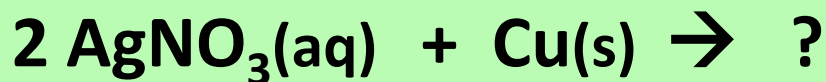
1. Writing equations for chemical reactions
2. Making and recording observations
3. The reactions we will do today
4. Procedure
5. Your lab report

Info for Introduction



1. Writing equations for chemical reactions

Here is a look at our first reaction today – Procedure A1, page 28. We use aqueous silver(I) nitrate and copper metal. A color change is a definite sign that a reaction took place.

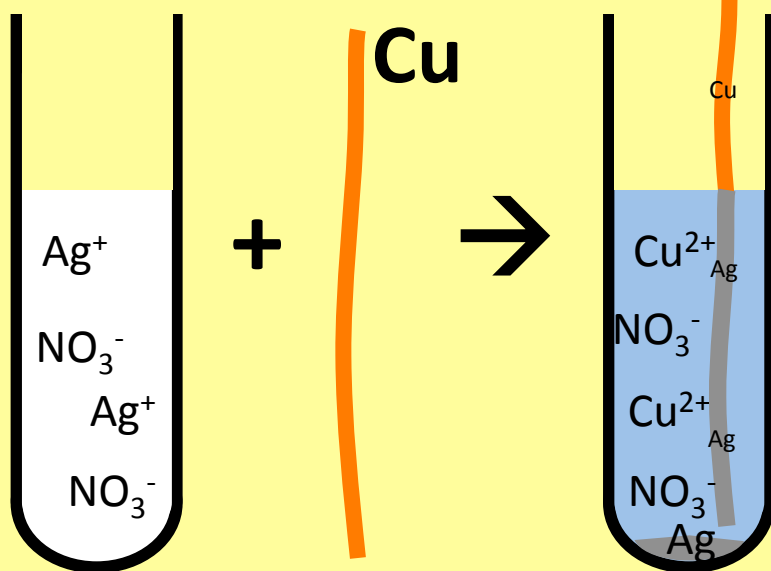


Remember –
All ionic salts
that dissolve
dissociate
100% into
ions in water.

The copper wire turns gray or black – elemental silver and the solution turns a bit blue – a sign of Cu^{2+} ... I like blue..

1. Writing equations for chemical reactions

Here is the balanced equation for the reaction.

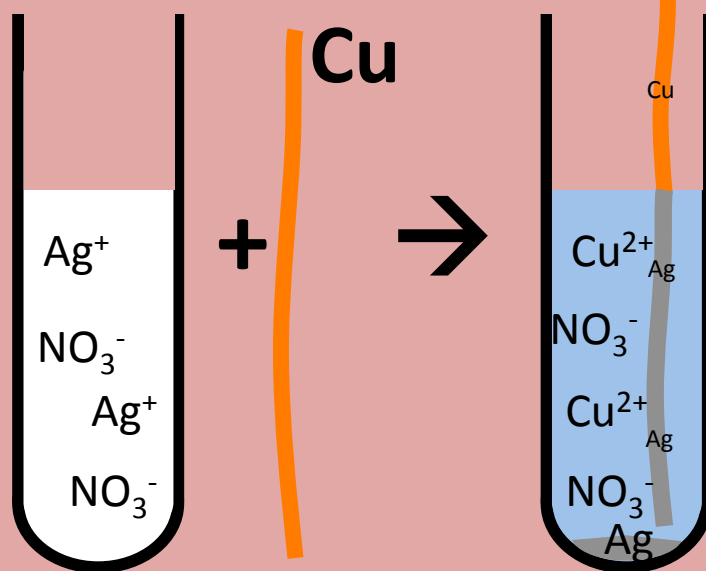
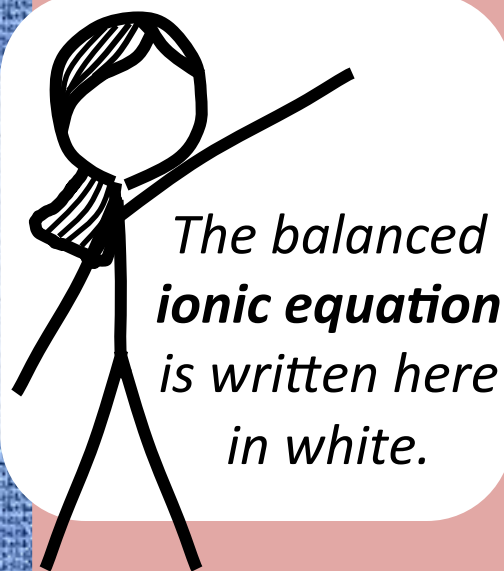


*This is a redox reaction!
Silver was reduced and copper was oxidized.
Silver was the oxidizing agent and...*

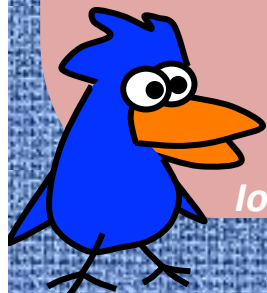


States of matter matter.

1. Writing equations for chemical reactions

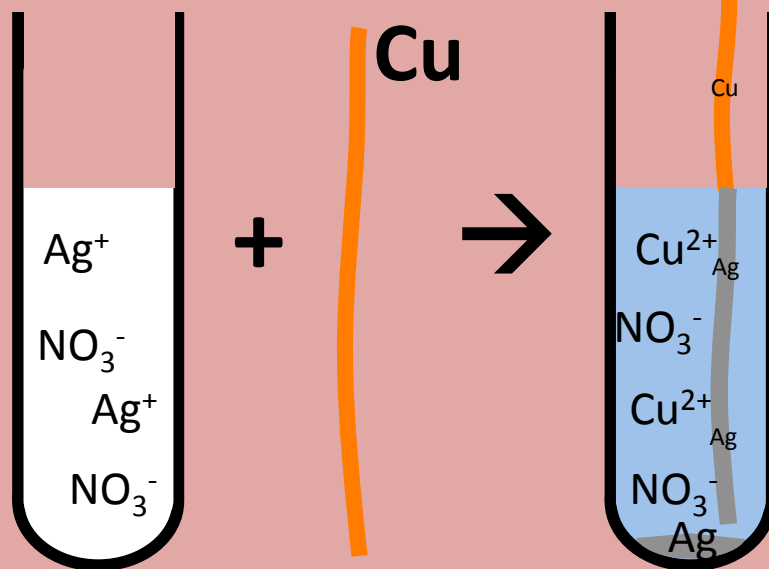
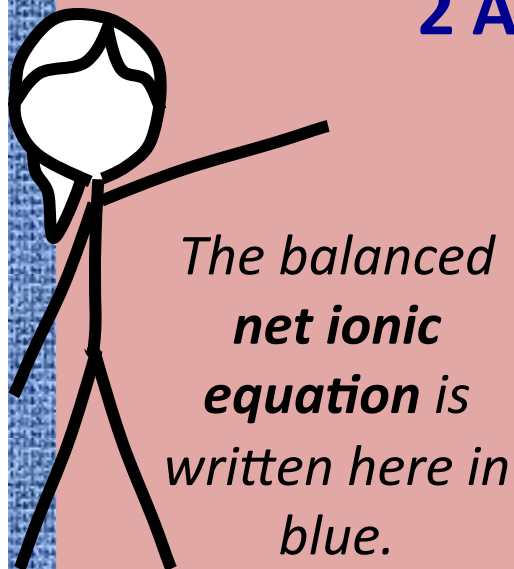


The cool thing about the **ionic equation** is that it tells us every detail in the reaction, omitting nothing.



Ion charges matter.

1. Writing equations for chemical reactions



Net ionic equations leave off the spectator ions.

Ahhh, that distinct smell of a quiz question.

2. Making and recording observations

Turned blue.

Within 10 s after the copper wire was added, the wire started darkening to gray. After a minute or so, the solution...

The day dawned chilly and cloudy. My lab manual has a curious stain from last week's experiment. One of our test tubes has a funny little chip out of it. My lab partner tipped over the silver nitrate. We...



Not good Very good Ummm... no.

3. The reactions we will do today – Part A2

Stuff burns. It's called combustion.

Combustion means combination with oxygen.

Hydrocarbons are compounds of hydrogen and carbon, and they all burn to produce carbon dioxide and water. Oxygen comes from the air.

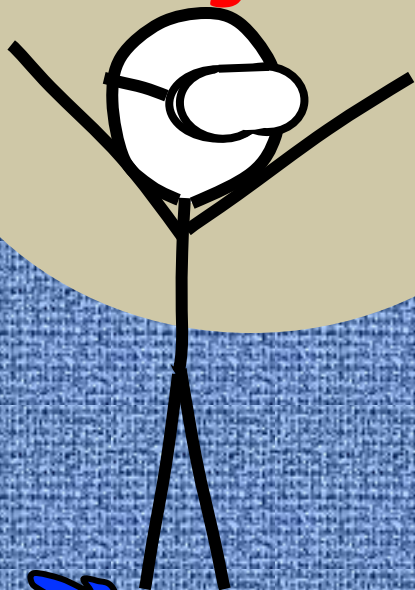
We will burn methane, CH₄(g) with our Bunsen burners. Methane is the main component of natural gas.



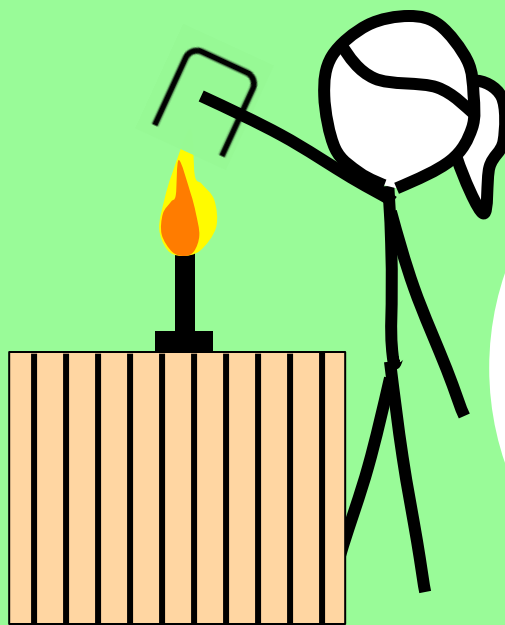
Nothing is ionic so no net ionic equation.

3. The reactions we will do today – Part A2

What are you doing???



Science. If I hold the beaker above the flame for a split second, it will collect carbon dioxide and water vapor.



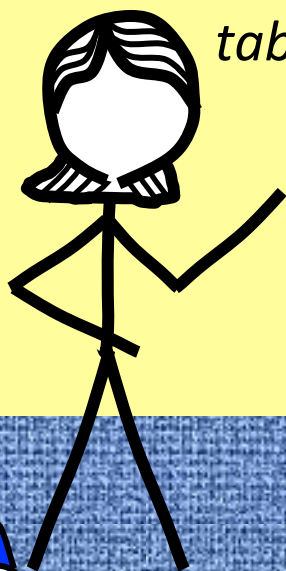
And the water vapor condenses –
 $H_2O(g) \rightarrow H_2O(l)$



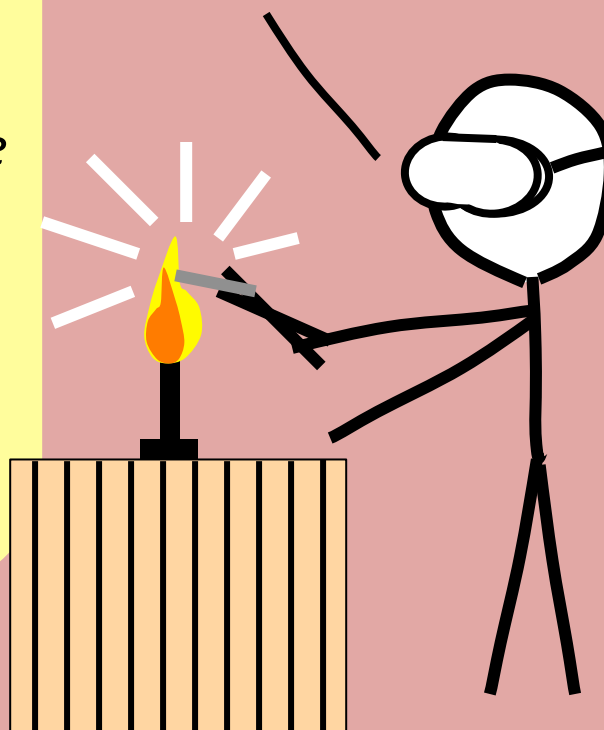
Nothing is ionic so no net ionic equation.

3. The reactions we will do today – Part A3

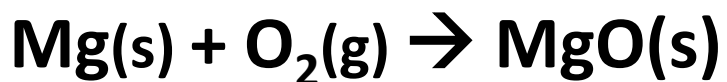
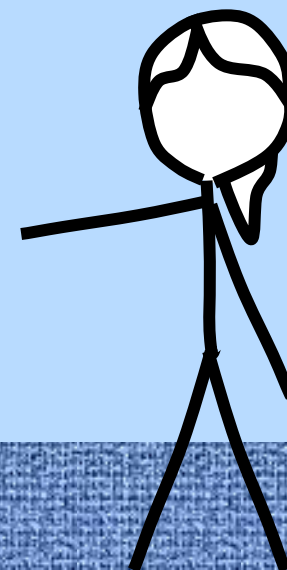
Instead of procedure A3 in the lab manual, Dr Mattson will give each group a 2.54 cm piece of magnesium to burn at the middle table.



Oooo. Bright.
Bright light!



When the $\text{MgO}(s)$ ash is cool, add it to a beaker and save it for Part C2.



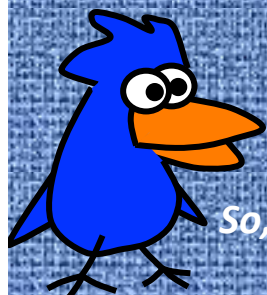
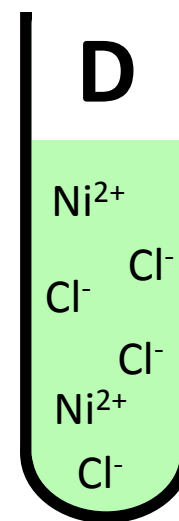
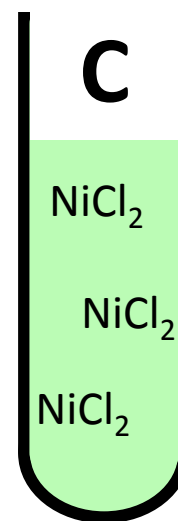
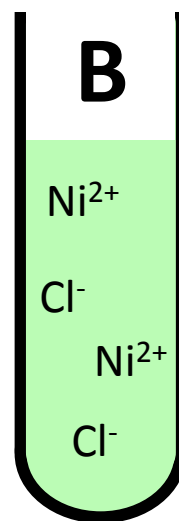
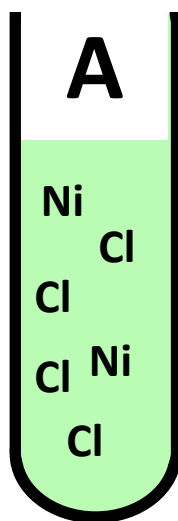
Nothing is dissolved in water
so no net ionic equation.

3. The reactions we will do today – Part B



In Part B, we will be using two solutions – aqueous nickel(II) chloride and aqueous sodium carbonate.

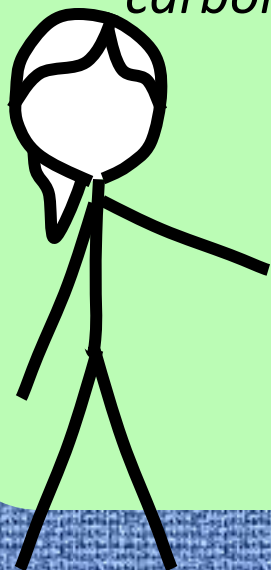
Which of these test tubes best represents aqueous nickel(II) chloride?



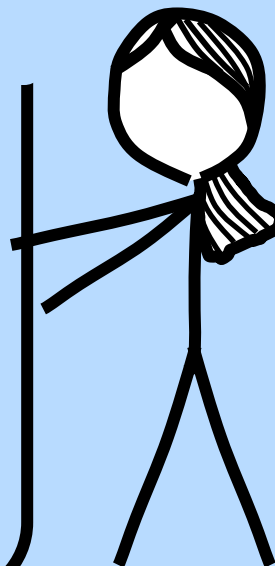
So, what is the answer?

3. The reactions we will do today – Part B

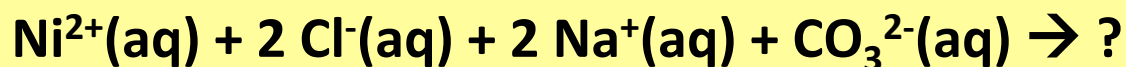
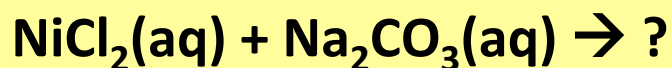
Sketch a solution of aqueous sodium carbonate.



Follow the lab manual procedure to add the two solutions in a little beaker. Record all observations.



It's a precipitate!
Let's complete these three equations!
Overall, ionic and net ionic.




Oooo! Oooo! Ions in aqueous solution!

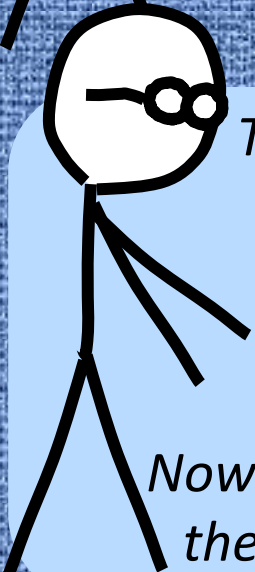
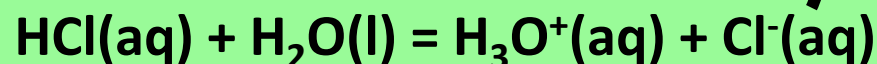
3. The reactions we will do today – Part C1



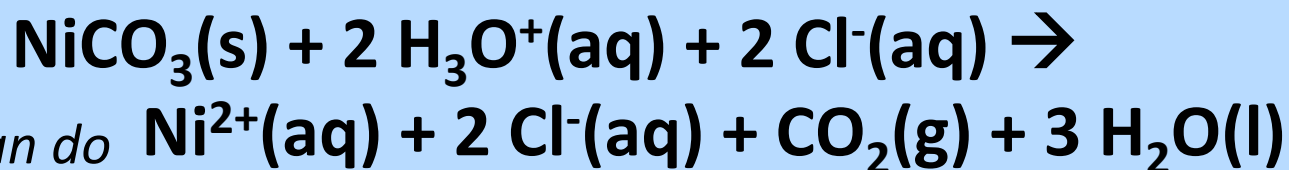
We're on to Part C1 now – acid reactions. This is the overall equation



Next we want to write the ionic equation. Solids are kept together, such as $\text{NiCO}_3(\text{s})$. So are covalent molecules such as $\text{CO}_2(\text{g})$. Strong acids dissociate and are written as ions. Instead of $\text{H}^+(\text{aq})$, we prefer to write it as $\text{H}_3\text{O}^+(\text{aq})$.

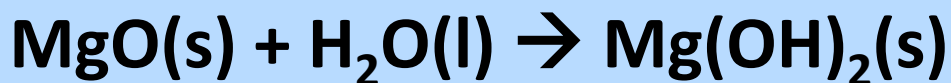


This is the ionic equation – crazy, 'eh?




Now you can do the net ionic.

3. The reactions we will do today – Part C2



Lastly, we do Part C2. Magnesium oxide soaks up water moisture from the air, but that takes a long time. We can speed it up.



This type of reaction is called hydrolysis.

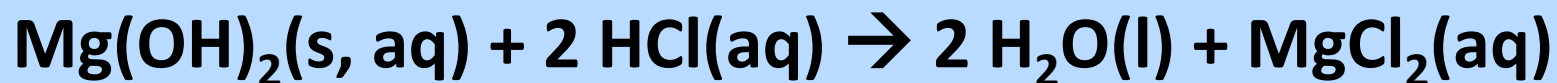


A tiny bit of magnesium hydroxide dissolves – enough to turn phenolphthalein pink to indicate the presence of hydroxide. Sooo, magnesium hydroxide is a base!



He has his goggles on again

3. The reactions we will do today – Part C2



This is the overall equation for the acid-base reaction. Water is always a product of an acid-base reaction.

We can write $\text{Mg(OH)}_2(\text{s, aq})$ because it is mostly insoluble, but a little dissolves as we just saw.

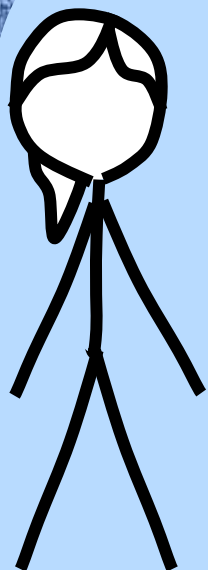
You can do the ionic equation. Leave $\text{Mg(OH)}_2(\text{s, aq})$ put together because it is mostly $\text{Mg(OH)}_2(\text{s})$. Remember $\text{HCl}(\text{aq})$ is a strong acid written as $\text{H}_3\text{O}^+(\text{aq}) + \text{Cl}^-(\text{aq})$. Water is covalent and stays together, but $\text{MgCl}_2(\text{aq})$ is ionic.



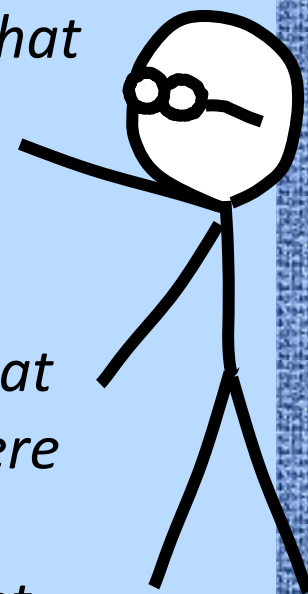
And then do the net ionic equation.



4. Procedure



- I. *Wear your safety glasses today. And we need Level 1b-i-ii attire today (Dress for a mess).*
- II. *Follow the procedures as written, except that we will be doing Part A3 ourselves. Dr Mattson will provide the piece of magnesium.*
- III. *The cover sheet summarizes everything that you need to include with your report. There is an emphasis on observations and balanced reactions, including ionic and net ionic when appropriate.*



5. Your lab report.

- ① First, the cover page with TA initials.
- ② Next, the trimmed copy pages from your lab notebook stapled together.
- ③ Turned in lab report *today* or *before* the start of class tomorrow.
- ④ Oh... and no cartwheels in lab.

Stick people inspired by xkcd
cartoons by Randall Munroe
(www.xkcd.com)

