

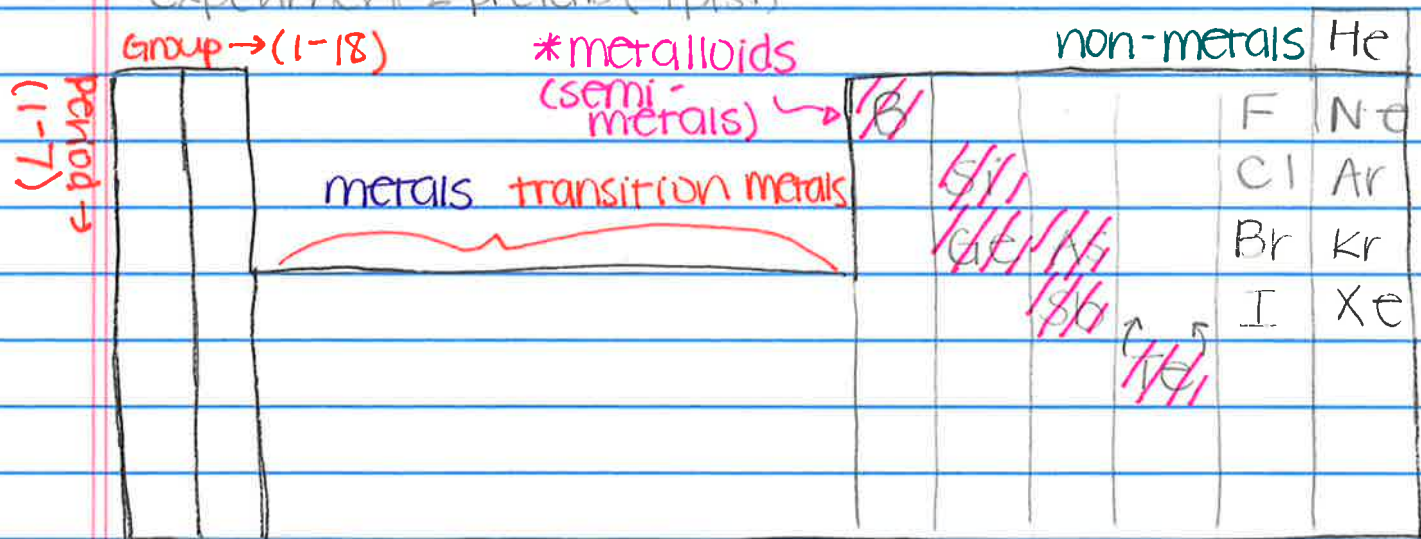
Today → start ch. 2

August 28<sup>th</sup>

Thursday → problem club w/ Kendall (7:30-9pm) Eppley III

Friday → sections 2.10-2.12 → nomenclature (have ions memorized by Friday)

Tuesday 9/3 → Expt. 2 → quiz over experiment 1 (6 pts.) + experiment 2 prelab (4 pts.)



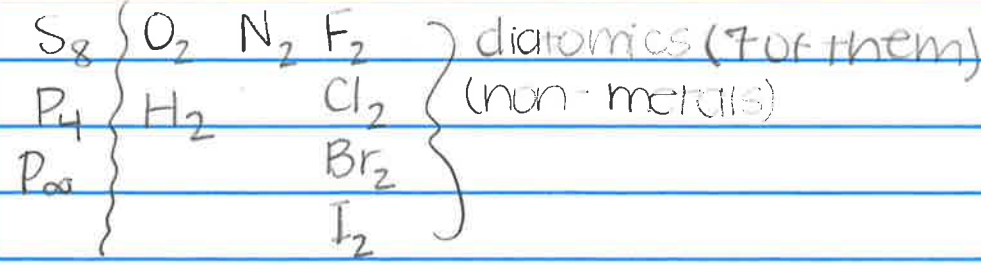
↑ alkali metals      ↑ alkaline earth metals      ↑ halogens      ↑ noble gases

\*main group elements

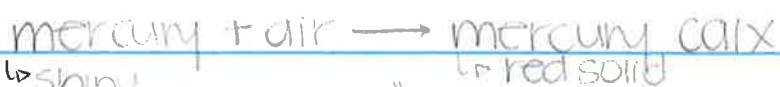
lanthanides (metals)  
actinides (metals)

metals ← malleable  
          ← conduct electricity  
          ← shiny  
          ← ductile

In standard conditions:

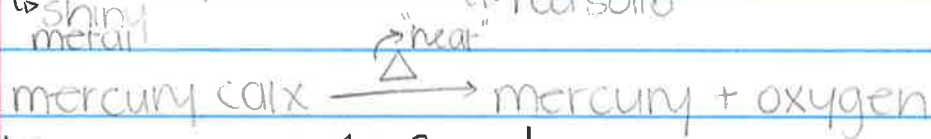


August 28<sup>th</sup>



↳ shiny metal

↳ red solid



5.00 g  $\rightarrow$  4.63 g

92.6% mercury & 7.4% oxygen

▷ take any sample + any size + think in terms of %.

↳ \* Law of constant composition: if you make something that is pure, the % composition doesn't change

Make compounds with copper and oxygen

CuO ↳ compound 1 79.9% copper + 20.1% oxygen

Cu<sub>2</sub>O compound 2 88.8% copper + 11.2% oxygen

Cu  $\rightarrow \frac{79.9}{20.1} = 3.975$

O

$\frac{88.8}{11.2} = 7.929$

$\frac{7.929}{3.975} = 1.99$

$\rightarrow \frac{2}{1}$  → ratio of simple whole #'s

↳ \* Law of multiple proportions

↳ CO<sub>2</sub> CO

H<sub>2</sub>O H<sub>2</sub>O<sub>2</sub>

CH<sub>4</sub> C<sub>3</sub>H<sub>8</sub>

Two compounds of boron and hydrogen

↳ compound A: 3.60g B for every 1g of H

↳ compound B: 7.2g B for 1.00g of H

$\frac{3.6}{1} = 3.6$

$\frac{7.2}{3.6} = \frac{2}{1}$

$\frac{7.2}{1} = 7.2$

↳