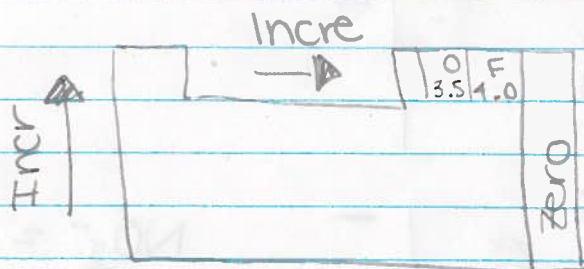
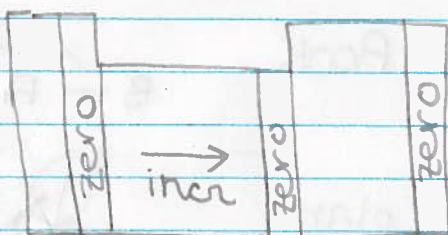


Electronegativity = how much an atom wants another e^-



- * "appetite" for an e^-
- * no units

Electron Affinity $E + e^- = E^-$



$$E_{mol} = \ominus \text{ KJ/mol}$$

↑
negative

- * "after" dinner

Steps for drawing Lewis dot structures

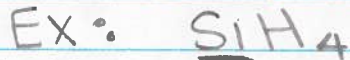
1) Sketch the atoms with Lewis dots



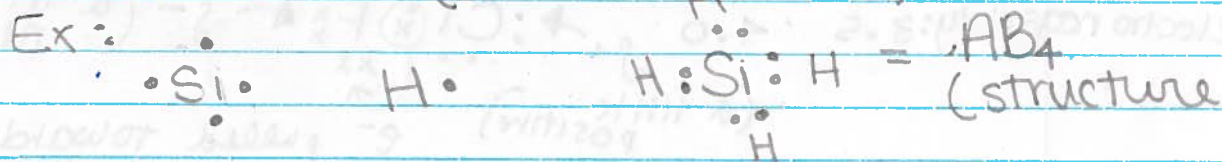
2) Adjust central atom for charge

- "A" = central atom
- "B" = bonded group
- "E" = e^- group

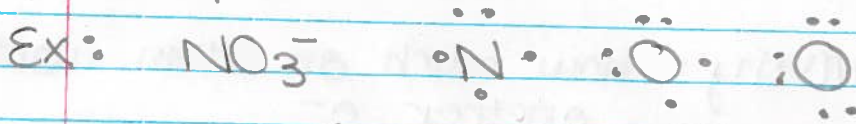
→ central atom is 1st atom listed



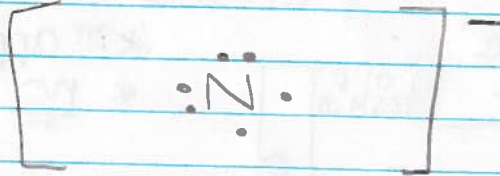
3) Add bonding groups ONE AT A TIME while forcing added group to obey octet rule (duet rule for H)



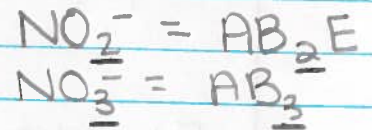
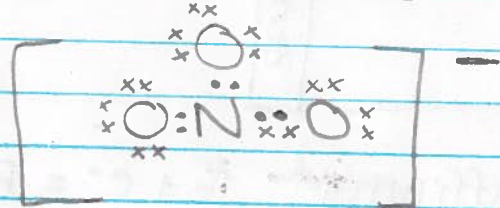
Step 1



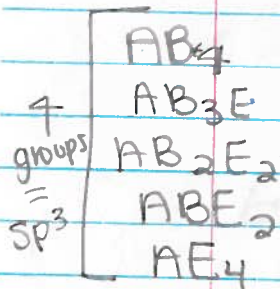
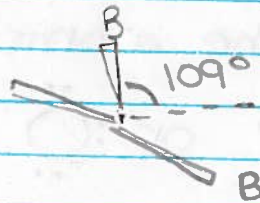
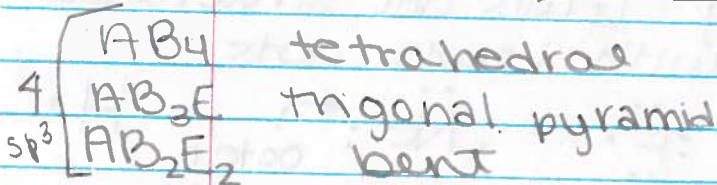
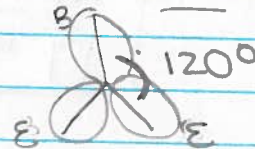
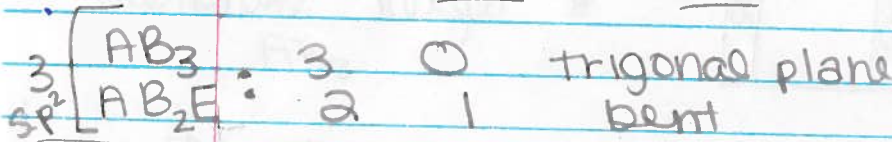
Step 2



Step 3



On table 8.1 in Book



mathematically mixing

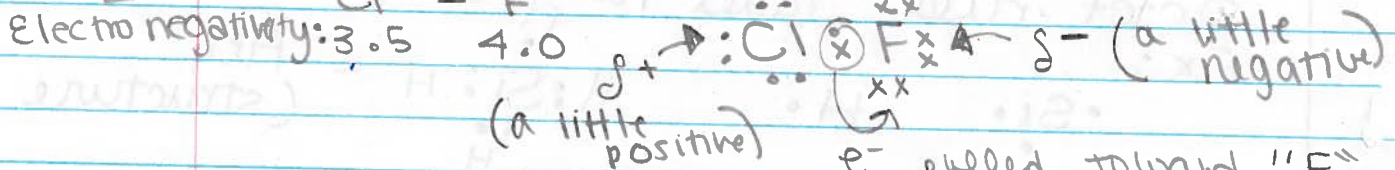
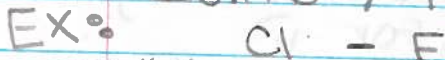
3 : p-orbitals

2 : 1s-orbitals

= "sp³"

orbitals

Ionic / POLAR / Non polar?



e⁻ pulled toward "F"

because it has a higher electronegativity

* POLAR BOND
 * POLAR molecule

