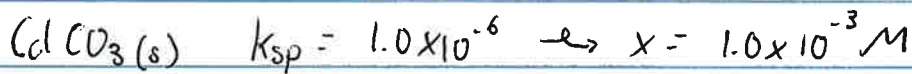
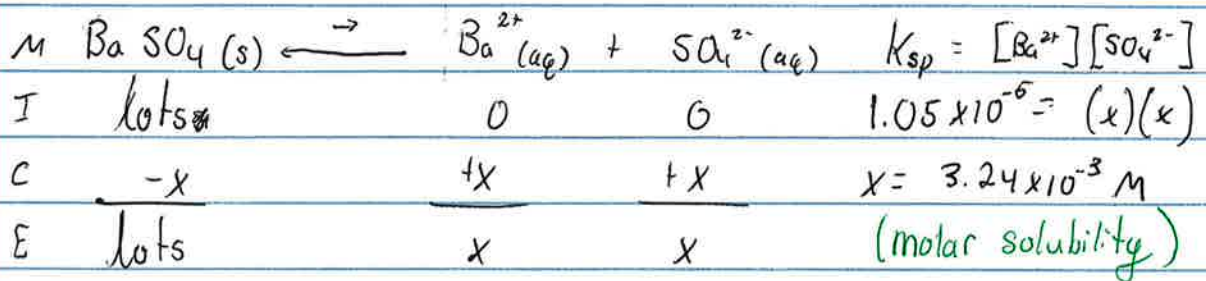
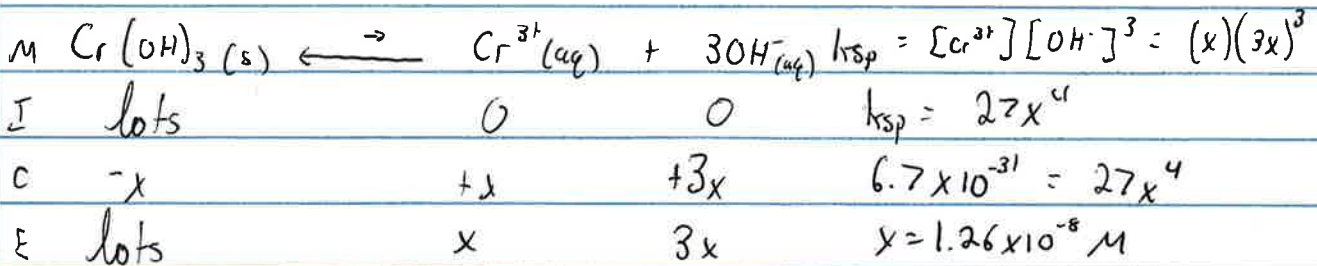
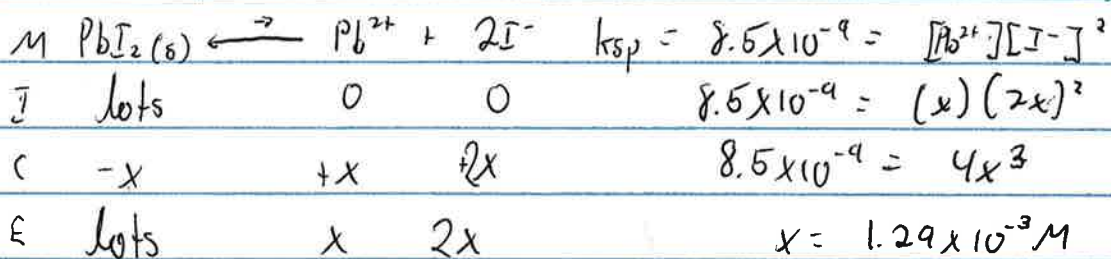
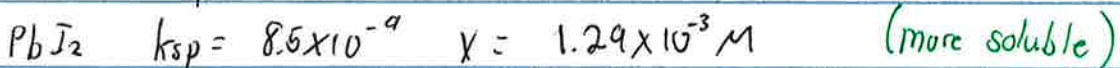
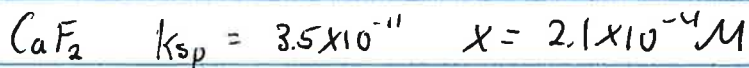


Today 3/14 (T-day) Finish ch. 16
 Thursday 3/15 Expt. 8 buffers 1
 Friday 3/16 Start ch. 17

*X will always be small

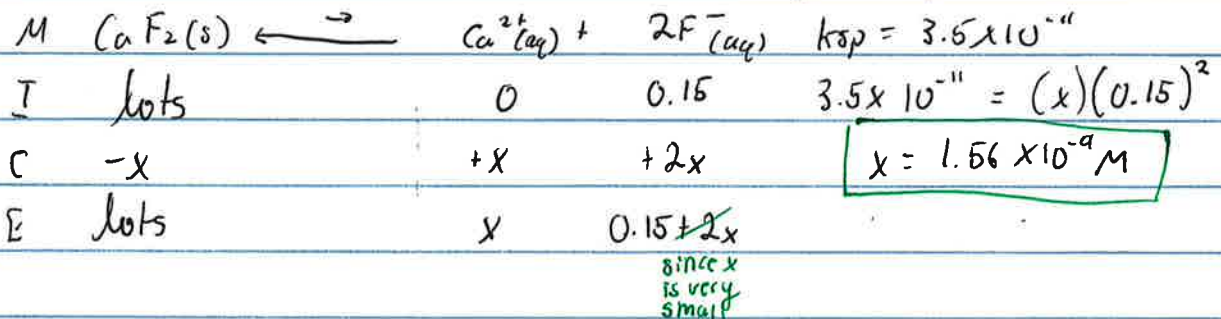


*when they have the same "salt ratio" (i.e. 1:1), you can directly compare K_{sp}



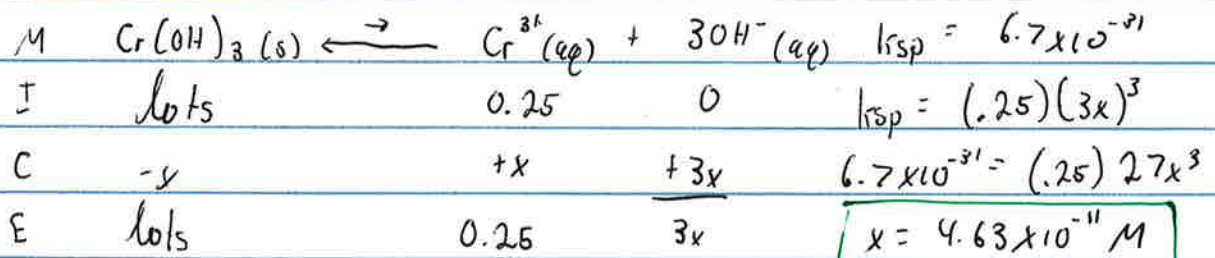
Common Ion

What is the molar solubility of CaF_2 in a solution that is 0.15 M NaF (makes Na^+ and F^-)? $K_{sp} = 3.5 \times 10^{-11}$

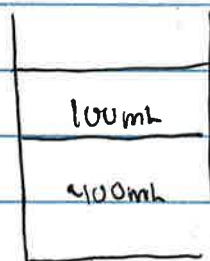
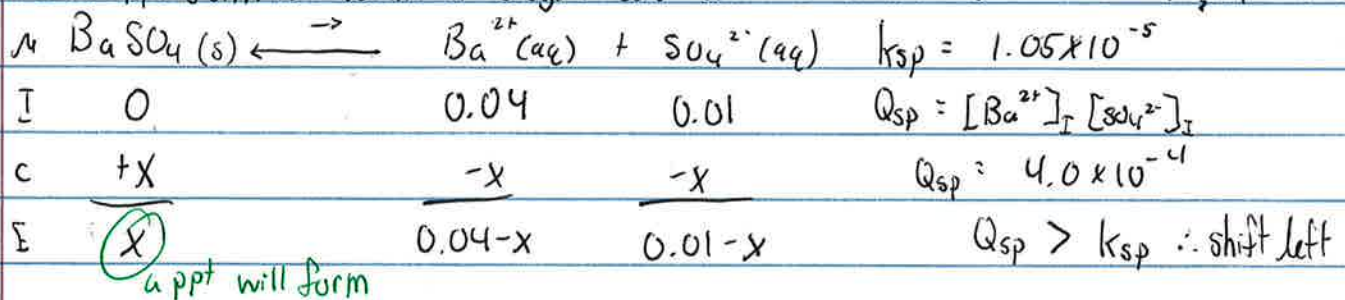


*calculator tip: Do not round until the very end

What is the molar solubility of $\text{Cr}(\text{OH})_3$ in a solution that is $0.25\text{ M Cr}(\text{NO}_3)_3$?



Will ppt form if 400 mL of $0.05\text{ M Ba}(\text{NO}_3)_2$ and 100 mL of $0.05\text{ M Na}_2\text{SO}_4$ are mixed?



0.050 M SO_4^{2-}
 0.050 M Ba^{2+}

$$M_c V_c = M_d V_d$$

$$(0.050\text{ M})(100\text{ mL}) = M_d (500\text{ mL})$$

$$M_d = 0.01\text{ M}$$

$$(0.050\text{ M})(400\text{ mL}) = M_d (500\text{ mL})$$

$$M_d = 0.040\text{ M}$$

