Today, Aug 29 : Sect 2.1 - 2.5
cut out flashcards and start memorizing
Friday, Aug 31 : Sect 2.10 - 2.12

· Conservation of matter experiment / definite composition:

cinnebar (red) $\xrightarrow{\text{heat}}$ mercury

5.00g Cinnebar → 4.63g Hg
7.00g Cinnebar → 6.48 Hg
27.00g Cinnebar → 25.00g Hg

} makes 93%

formula of Cinnebar : HgO

· another experiment:

color:

copper oxide #1 (brick)	copper oxide #2 (black)
79.9% Cu	88.8% Cu
20.1 % O	11.2 % O
$\frac{\%Cu}{\%O} = 3.98$	$\frac{\%Cu}{\%O} = 7.93$

$\frac{3.98}{7.93} = .502$ (circled)

you will end with a ratio of small whole numbers
(law of multiple proportions)
↳ lead to use of formulas

~~Compound 1~~
~~3.60g A~~
~~1.00g B~~

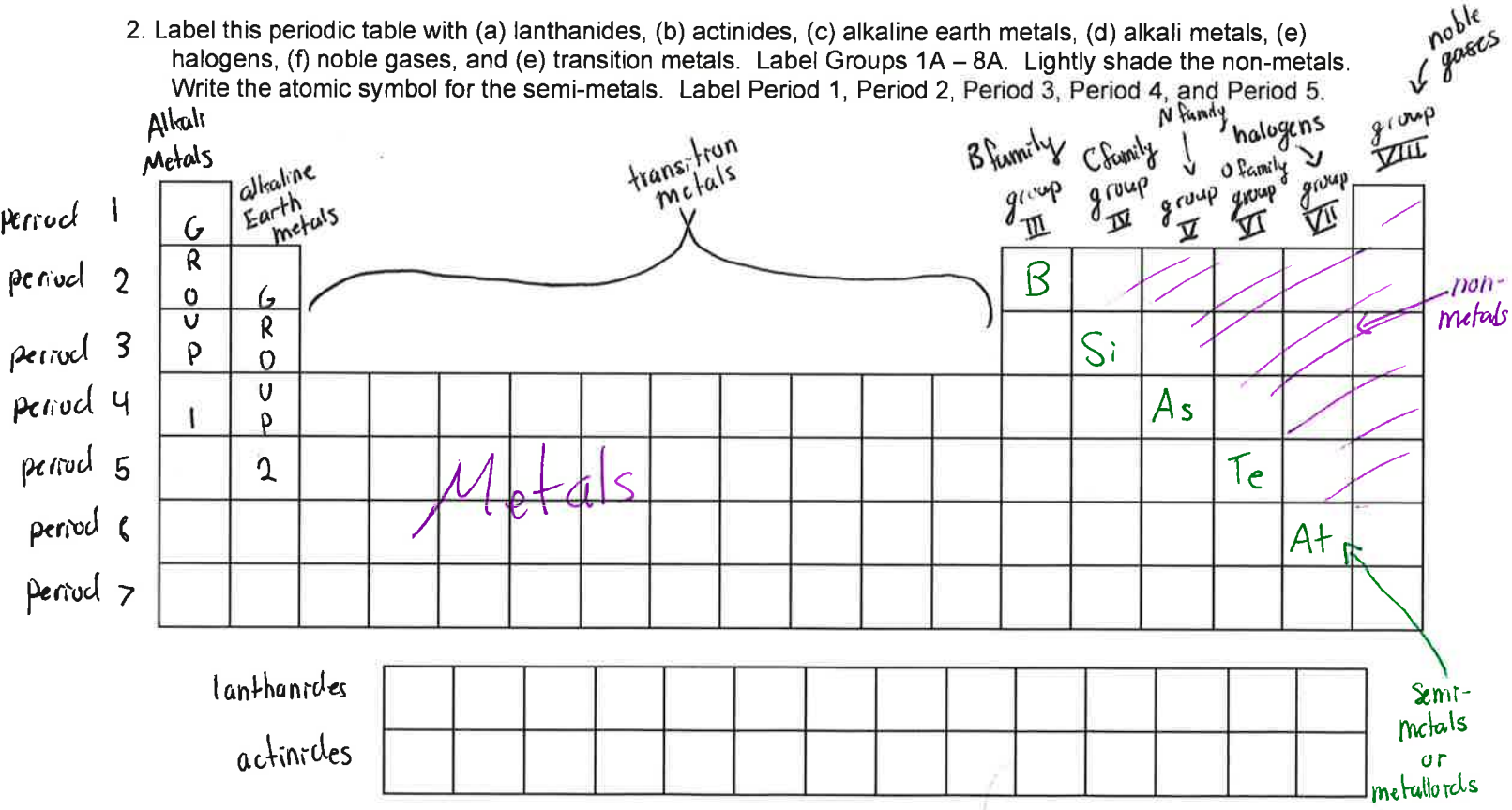
~~Compound 2~~

Iron experiment instead

1. These are the 51 elements for which you should know name and atomic symbol (not shown: $_{92}U$). Write the names of the elements with atomic symbols are NOT based on the first one or two letters of the element's name.

1 H																	2 He
3 Li	4 Be											5 B	6 C	7 N	8 O	9 F	10 Ne
11 Na	12 Mg											13 Al	14 Si	15 P	16 S	17 Cl	18 Ar
19 K	20 Ca	21 Sc	22 Ti	23 V	24 Cr	25 Mn	26 Fe	27 Co	28 Ni	29 Cu	30 Zn	31 Ga	32 Ge	33 As	34 Se	35 Br	36 Kr
37 Rb	38 Sr				42 Mo				46 Pd	47 Ag	48 Cd		50 Sn	51 Sb		53 I	54 Xe
55 Cs	56 Ba				74 W				78 Pt	79 Au	80 Hg		82 Pb	83 Bi			86 Rn
87 Fr	88 Ra																

2. Label this periodic table with (a) lanthanides, (b) actinides, (c) alkaline earth metals, (d) alkali metals, (e) halogens, (f) noble gases, and (e) transition metals. Label Groups 1A – 8A. Lightly shade the non-metals. Write the atomic symbol for the semi-metals. Label Period 1, Period 2, Period 3, Period 4, and Period 5.



*families/groups tend to share physical ~~other~~ properties and chemical properties