

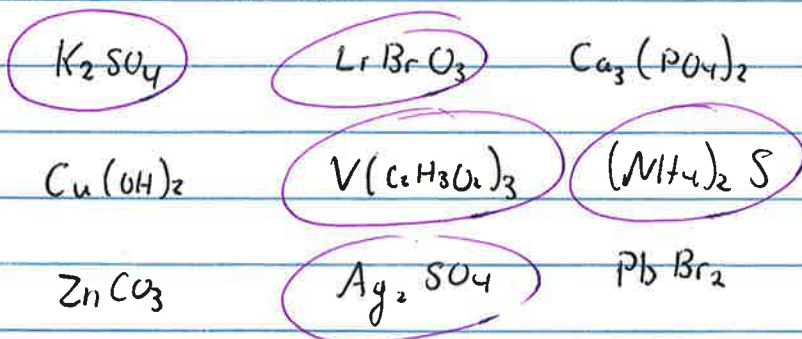
Today sect 4.5-4.9

Sunday: Ali's problem club

Monday: Review and nom II

Tuesday: expt. 5

Which of these are soluble?

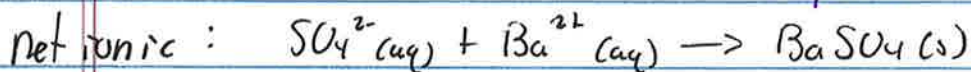
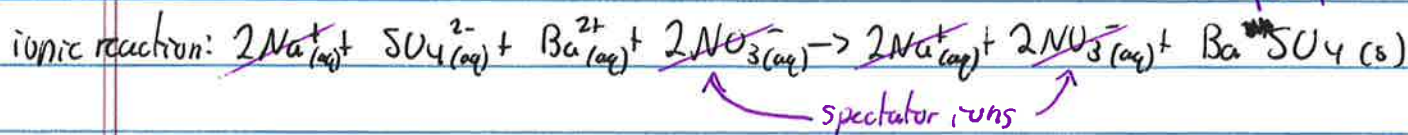
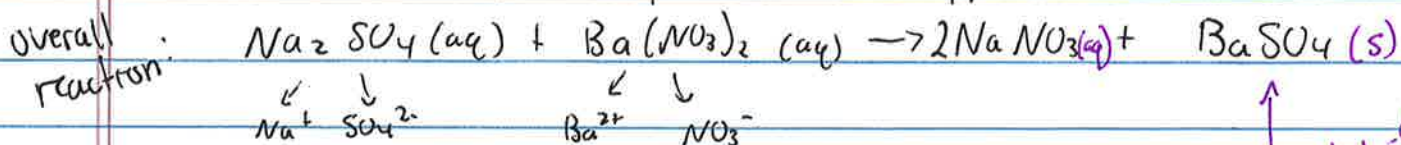


all ionics that ~~are~~ dissolve, dissociate 100% into ions in solution
↳ "strong electrolyte" = soluble ionic

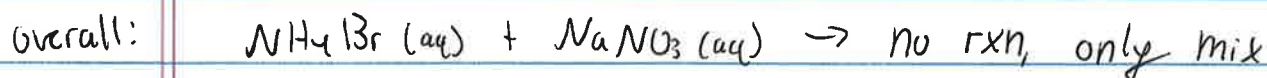
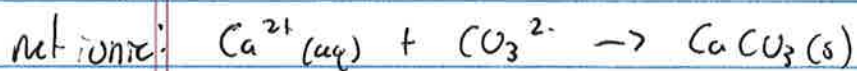
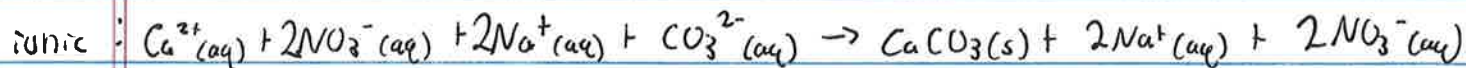
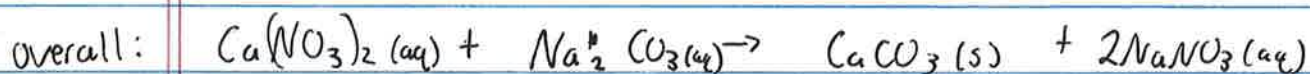
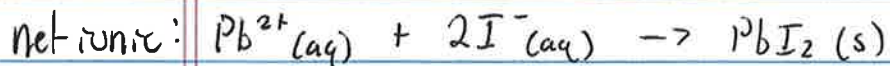
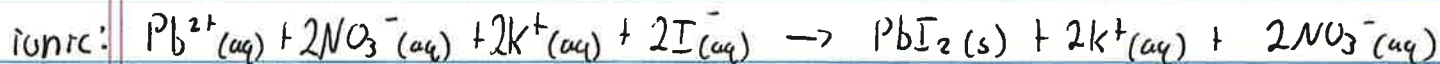
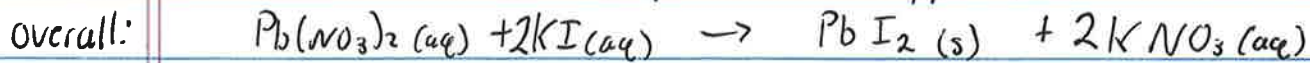
• Insoluble ionics are non-electrolytes

• Sparingly soluble ionics are weak electrolytes

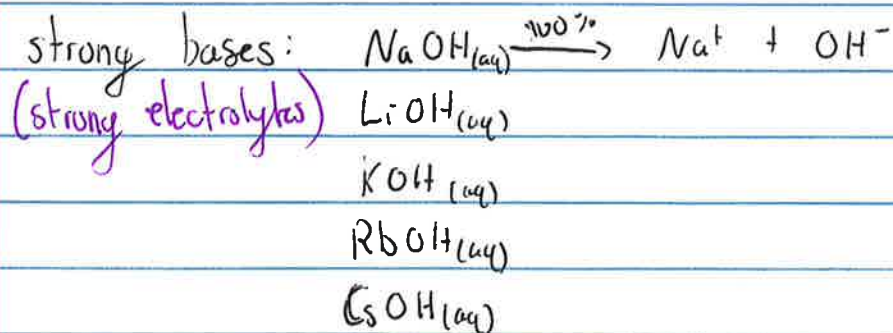
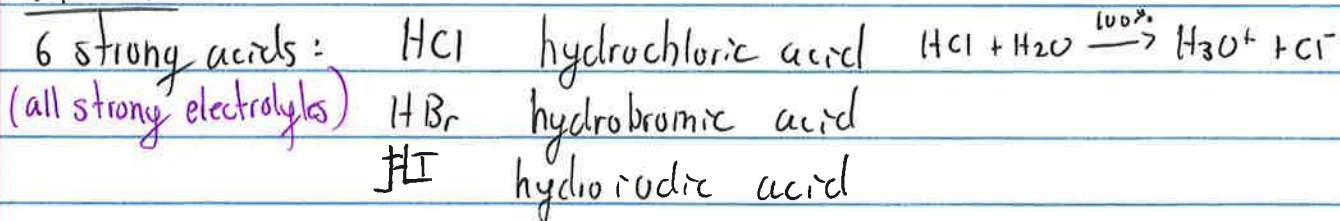
• Mix two solutions and predict if a ppt forms.



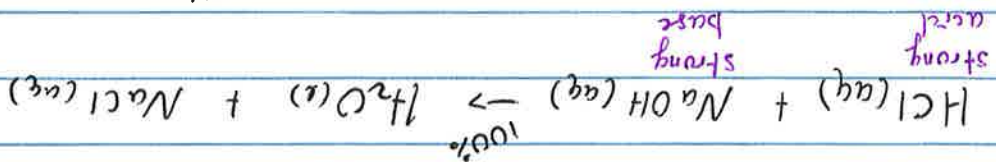
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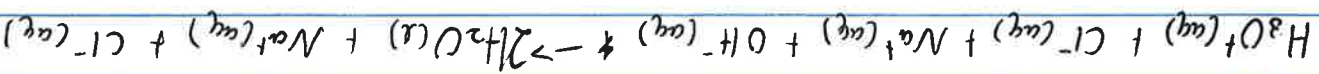
Acids



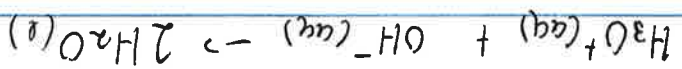
Overall:



Ionic:



Net Ionic:



weak acid = weak electrolyte

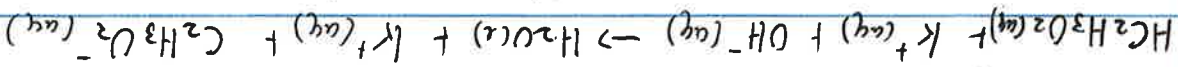


weak acid + strong base \rightarrow H₂O + salt

Overall:



Ionic:



Net Ionic:

