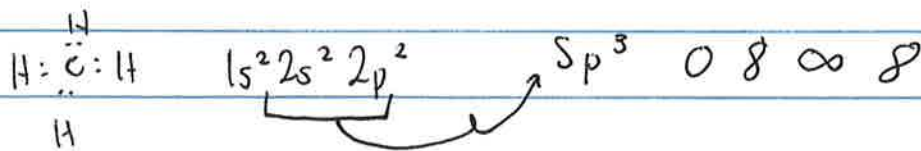
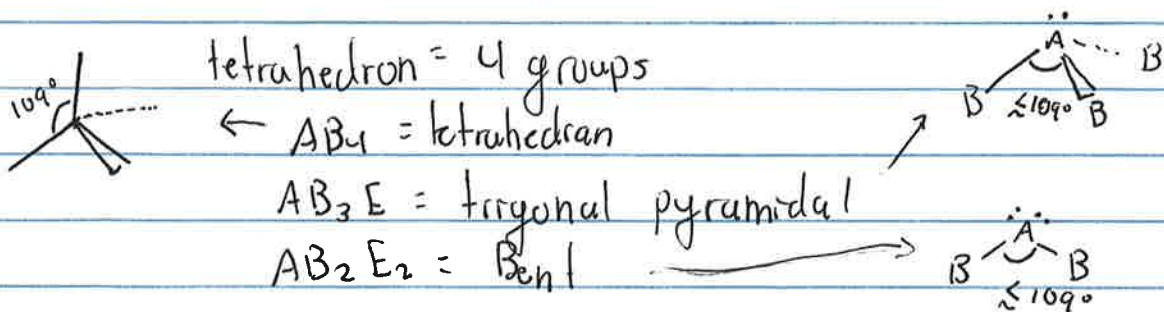
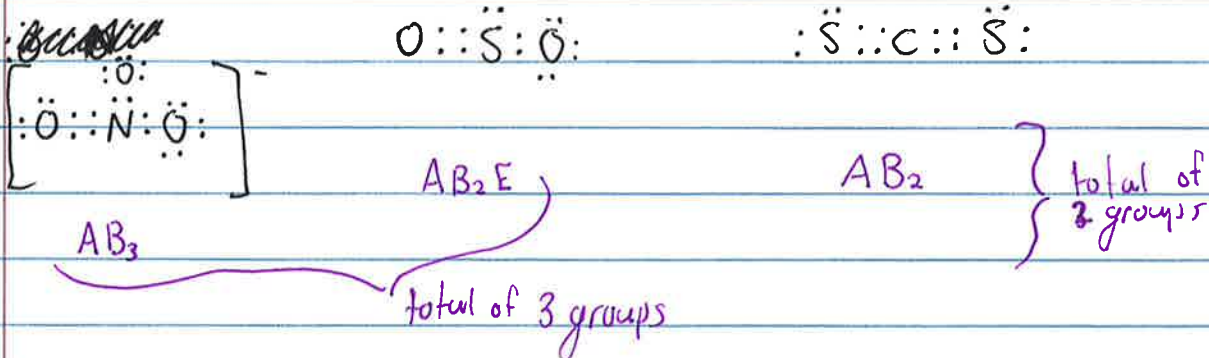
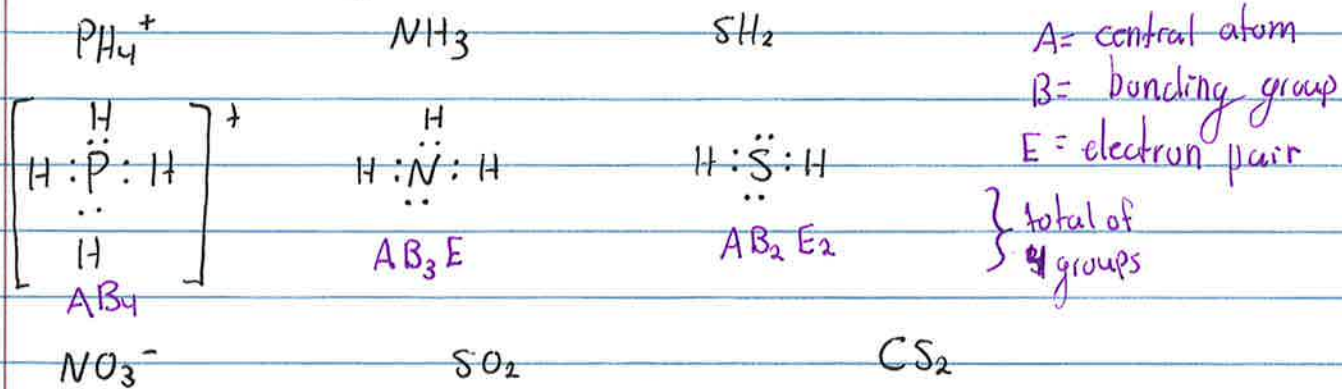


Today Oct 29: Continue with ch 7 and start ch 8  
 Lab tomorrow: molecular shape

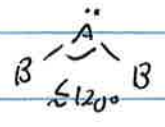
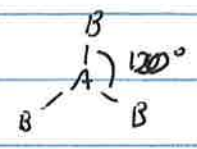
warm-up activity: sketch the lewis dot structure



3 groups (120°) (sp<sup>2</sup>)

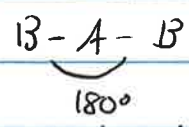
AB<sub>3</sub> = trigonal planar

AB<sub>2</sub>E = Bent



2 groups (180°) (sp)

AB<sub>2</sub> = linear



polar bond => polar molecule

$\delta^+$  H -  $\delta^-$  F

2.1 4.0

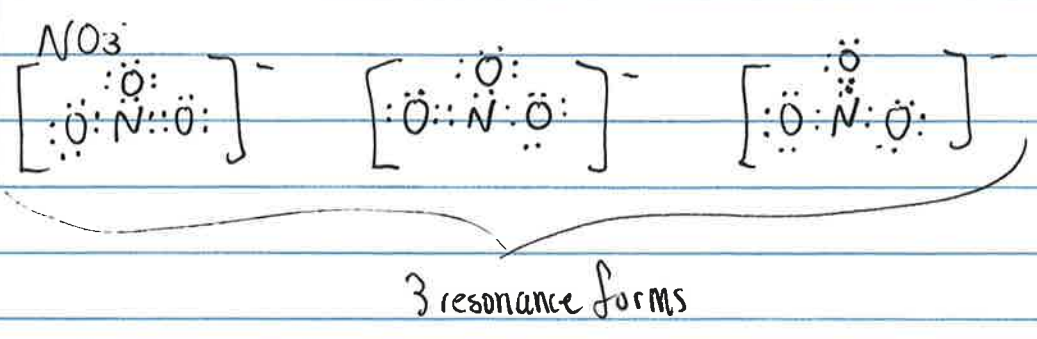
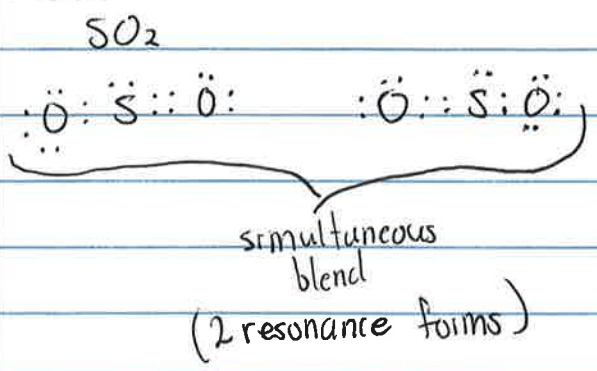
polar bonds - they cancel each other so non-polar molecule

$\ddot{\text{O}}::\text{C}::\ddot{\text{O}}:$

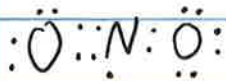
3.5 2.1 3.5

\* if there is an E group the molecule is polar

### Resonance

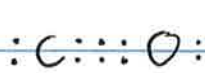


$\text{NO}_2$



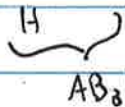
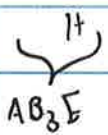
↑ unpaired electron = paramagnetic

$\text{CO}$



} one of O's electrons goes to C

2 central atoms



← each get their own AB<sub>3</sub>E formula