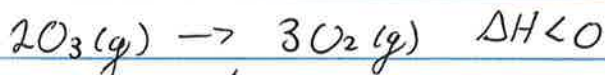


Today Feb 4: Catch-up and review

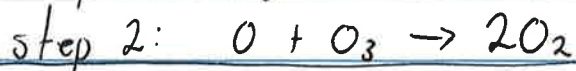
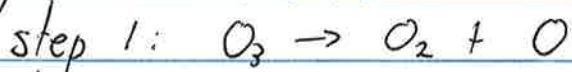
Tuesday: Expt 3, problem club with Ali: 6:30 - 7:30 eply 107

Wednesday: CK1, doors open at 10:45

Friday: start ch 14



proposed mechanism:



intermediate

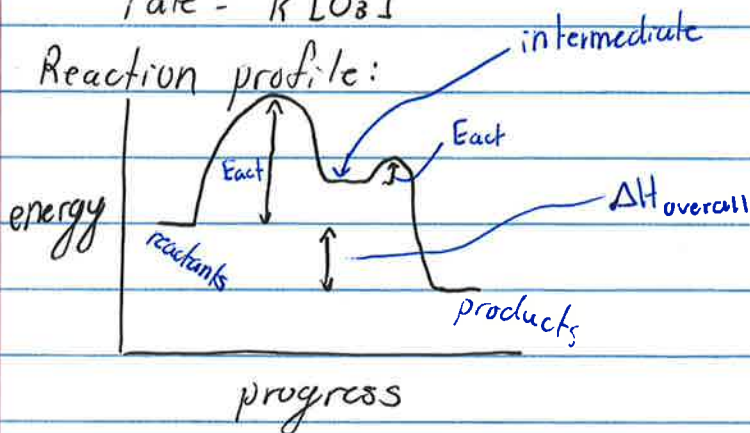
slow step

unimolecular step

bimolecular step

One can write the rate law from the reactants of the slow step in a mechanism.

$$\text{rate} = k[\text{O}_3]^2$$



un-related reaction profile



what we know:

- 2 steps
- second step slow
- endothermic
- has intermediate

