Part 1. Naming Ionic Compounds.

barium hydroxide	
strontium nitrite	
ammonium bromide	
potassium sulfate	
iron(II) phosphide	
iron(III) peroxide	
barium carbonate	
ammonium acetate	
lithium sulfite	
calcium dichromate	
tin(II) oxide	
beryllium phosphate	
ammonium chromate	
lithium permanganate	
calcium bicarbonate	
potassium sulfide	
ferrous cyanide	
ammonium nitrate	
iron(III) bisulfite	
ferric chloride	

	$LiNO_2$
	$Cu(ClO_4)_2$
	Na_2SO_4
	NH_4NO_3
	$NiCl_2$
	$NiCrO_4$
	$Au(HSO_4)_3$
	BaS
	$(NH_4)_2O_2$
	$KMnO_4$
	$Cr_2(CO_3)_3$
	AgI
	$Ba(OH)_2$
	NH_4CN
	$Sn(ClO_3)_2$
	VBr_3
	$AgClO_4$
	$Ca_3(PO_4)_2$
	$Al_2(SO_4)_3$
	$NaC_2H_3O_2$

Part 2. Covalent Molecules.

sulfur dichloride	
diarsenic pentoxide	
sulfur tetrachloride	
carbon tetrafluoride	
trisulfur diphosphide	
phosphorus trichloride	
dinitrogen pentasulfide	
selenium dibromide	
sulfur hexafluoride	
selenium monosulfide	
xenon tetrafluoride	
carbon monoxide	
iodine monochloride	

	CO_2
	BrF_3
	N_2O_3
	NO_2
	OCl_2
	SiO_2
	IBr_3
	As_2O_5
	N_2O
	ClO_2
	CO_2
	PCl_3
	$SiCl_4$

arsenic trifluoride	
tellurium trioxide	
ammonia	
carbon disulfide	
diphosphorus pentoxide	
diarsenic trioxide	
boron trifluoride	

	P_4S_6
	BrF_5
	CO
	P_4S_{10}
	SF_2
	IF_7
	H_2S

Part 3. Acids.

nitrous acid	
sulfuric acid	
hydrobromic acid	
acetic acid	
nitric acid	
sulfurous acid	
chlorous acid	
hydroiodic acid	
hydrofluoric acid	
hypobromous acid	
periodic acid	
bromic acid	
phosphoric acid	
hydrochloric acid	
chloric acid	
hydroiodic acid	
hypochlorous acid	
bromous acid	
perchloric acid	
hydrobromic acid	

	HCl
	HClO
	$HClO_2$
	$HClO_3$
	$HClO_4$
	H_3PO_4
	HBr
	HI
	HF
	$HC_2H_3O_2$
	HNO_2
	H_2SO_4
	H_2SO_3
	$HBrO_4$
	HNO_3
	HBrO
	$HBrO_2$
	$HBrO_3$
	$HBrO_4$
	HIO_3